

Group Name:

Lewis Dot Structure Group Lab Report

LibreTexts page: [7: Molecular Structure](#)

(<https://chem.libretexts.org/link?214684>)

Please don't edit, rearrange or delete anything that is already in this document. Just add your answers inside the boxes.

You can use shortcuts for superscripts and subscripts when needed:

X² Superscript Ctrl+.

X₂ Subscript Ctrl+,

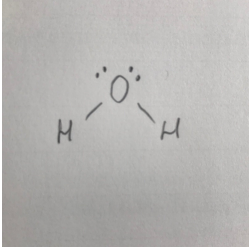
For each compound*:

1. Draw Lewis Dot Structure and any resonance structures. Take a picture of it and insert it in the table. Resize as needed.
2. Calculate formal charges of atoms. Show calculations for partial credit.

Hint: go to <https://chem.libretexts.org/link?52857>

*Use water as an example.

Compound	Lewis Dot Structure, Resonance structures	Formal charge of atoms
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H_2O		$\text{O} = 6 - 4 - (4/2) = 0$ $\text{H} = 1 - 0 - (2/2) = 0$
NH_3		
NH_4^+		
SO_2		
NO_3^-		

SeCl_4		
AlCl_3		
XeCl_4		

